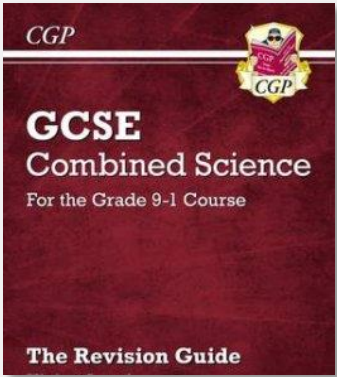






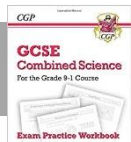
Step 1 What do I need to know?

- Use a reactivity series to predict the result of displacement reactions.
- Link reactivity to the number of energy shells and the distance of electrons from the nucleus.

Step 2 How do I find out about it?

Revision Guide Page		Web Links
		 Reactivity Series  Reactivity Series
Higher	Pg. 132	
Foundation	Pg. 130	
Triple (chemistry)	Pg. 133	

Step 3 What can I do to help me learn it?



Complete the relevant questions in your CGP Science Workbook

Higher	Pages 134
Foundation	Pages 115

TASK 1- Write balanced equations and word equations for the following reactions

1. Zinc and copper sulphate
2. Magnesium and silver nitrate
3. Potassium and zinc nitrate
4. Gold and copper sulphate

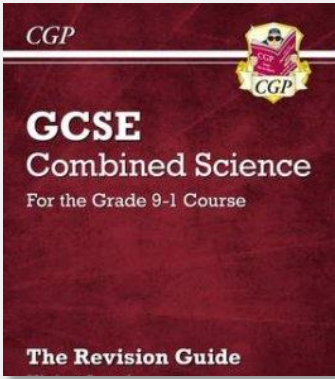


TASK 2- Describe what you would observe in the reaction of iron and copper sulfate



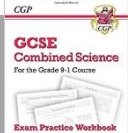
Step 1 What do I need to know?

- Describe the process of smelting and suggest metals which could be smelted
- Link ideas and recycling to the and further extracting to economic and environmental issues.

Step 2 How do I find out about it?

Revision Guide Page		Web Links
		 Metal Extraction Electrolysis  Extracting metals
Higher	Pg. 133-135	
Foundation	Pg. 130-132	
Triple (chemistry)	Pg. 137-144	

Step 3 What can I do to help me learn it?

	Complete the relevant questions in your CGP Science Workbook	Higher	Pages 136-138
		Foundation	Pages 115-117

TASK 1- Describe using equations how iron is extracted from iron ore in the blast furnace.

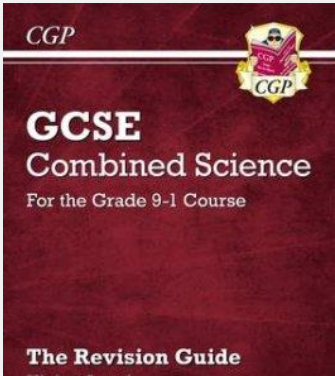
TASK 2- Describe how copper is purified using electrolysis.



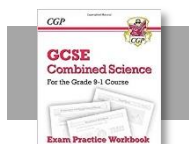
Step 1 What do I need to know?

- The definition of the term 'oxidation' and 'reduction.'
- The order of the reactivity series.
- How metals can be extracted using carbon.
- How iron is extracted using a blast furnace.
- How aluminium is extracted using electrolysis.
- Oxidation and reduction in terms of electron transfer.

Step 2 How do I find out about it?

Revision Guide Page		Web Links
		<p>Bitesize</p> <p>Metal extraction Oxidation & Reduction</p> <p>YouTube</p> <p>Metal extraction</p> <p>Oxidation & Reduction in Metal Extraction</p> <p>Oxidation and Reduction in Terms of Electrons</p> <p>Aluminium Extraction</p>
Higher	Pg. 133-4	
Foundation	Pg. 130-133	

Step 3 What can I do to help me learn it?



Complete the relevant questions in your CGP Science Workbook

Higher

Pages 137

Foundation

Pages 115

TASK 1- Define the terms 'oxidation' and 'reduction'.

TASK 2- Fill in the gaps below

When metals react with oxygen in the air they produce metal This is an example of an oxidation reaction because oxygen is to the metal. Some metals that are more reactive such as Others such as are very unreactive and never react with oxygen.

oxides added potassium gold

TASK 3- Write a paragraph describing iron extraction in terms of oxidation and reduction.

TASK 4- HT Only- Write balanced equations for the following reactions:

Magnesium + Oxygen → Magnesium Oxide

Iron + Oxygen → Iron Oxide

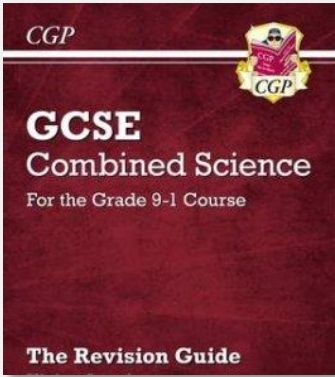
Copper + Oxygen → Copper Oxide



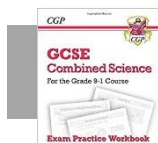
Step 1 What do I need to know?

- Identify acid, alkali and neutral on the pH Scale.
- Know the pH of common household chemicals.
- Know a variety of indicators and their limitations.
- Write a method for carrying out a neutralisation reaction.
- Write a balanced equation for a neutralisation reaction.

Step 2 How do I find out about it?

Revision Guide Page		Web Links
		<p>Bitesize</p> <p>pH and neutralisation reaction.</p> <p>YouTube</p> <p>Neutralisation</p> <p>titration experiment</p> <p>Measuring pH</p>
Higher	Pg. 131	
Foundation	Pg. 128-9	

Step 3 What can I do to help me learn it?



Complete the relevant questions in your CGP Science Workbook

Higher	Pages 130-2
Foundation	Pages 112-3

TASK 1- Complete the equation: **Acid + Alkali** → _____ + _____

TASK 2- What ion is produced when an acid is in solution?

TASK 3- What ion is produced when an alkali is in solution?

TASK 4- Write an equation for the formation of water in a neutralisation reaction.

TASK 5- Write a word and symbol equation for the reaction of potassium hydroxide and hydrochloric acid.

TASK 6- Explain, using ions, why water is formed during a neutralisation reaction

TASK 7- Evaluate the benefits of using a pH probe over universal indicator.



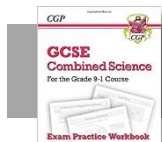
Step 1 What do I need to know?

- Know the definitions of 'acid', 'Insoluble base' and 'salt'.
- Describe the properties of insoluble bases.
- Explain how to produce a salt from an acid and insoluble base.

Step 2 How do I find out about it?

Revision Guide Page		Web Links
		 Acids and Bases Reactions with acids Copper Sulphate Practical
Higher	Pg. 129	
Foundation	Pg. 129	

Step 3 What can I do to help me learn it?



Complete the relevant questions in your CGP Science Workbook

Higher	Pages 132
Foundation	Pages 113

TASK 1- Define a base and an alkali.

TASK 2- Complete the equation: ACID + BASE → S___ + W_____

TASK 3- Write a method for the following practical: To make Copper Sulphate from a reaction between sulfuric acid and copper oxide. The diagrams below will help.

